

LECTURE 26 OXIDATION/REDUCTION

I. Adding and Subtracting Half-Cell Reactions

II. Nernst Equation

- A. Example: Calculate ΔE at 25°C of a cell in which the concentration of Zn^{2+} ions is 0.10 M and Cu^{2+} is 0.0010 M.
- B. Example: $\text{Pb}^{2+}(\text{aq}) + \text{Zn}(\text{s}) \rightleftharpoons \text{Zn}^{2+}(\text{aq}) + \text{Pb}(\text{s})$ What is K at 25°C?
- C. Reduction of vitamin B₁₂